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Electrodeposition of WO₃ Nanoparticles for Sensing Applications

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Additional information is available at the end of the chapter

<http://dx.doi.org/10.5772/61216>

Abstract

The motivation of using metal oxides is mainly due to its charge storage capabilities, and electrocatalytic, electrochromic and photoelectrochemical properties. But comparing with bulk, nanostructured materials present several advantages related with the spatial confinement, large fraction of surface atoms, high surface energy, strong surface adsorption and increased surface to volume ratio, which greatly improves the performances of these materials. The deposition of this materials can be accomplished by a variety of physical and chemical techniques but nowadays, electrodeposited metal oxides are generally used in both laboratories and industries due to the flexibility to control structure and morphology of the oxide electrodes combined with a reduced cost. Tungsten oxide (WO₃) is a well-studied semiconductor and is used for several applications as chromogenic material, sensor and catalyst. The major important features is its low cost and availability, improved stability, easy morphologic and structural control of the nanostructures, reversible change of conductivity, high sensitivity, selectivity and biocompatibility. For the electrodeposition of WO₃, more than one method can be adopted: electrodeposition from a precursor solution, anodic oxidation, and electrodeposition of already produced nanoparticles; however, in this case the mechanism of the electrodeposition is not fully understood. In this chapter, a review of the latest published work of electrodeposited nanostructured metal oxides is provided to the reader, with a more detailed explanation of WO₃ material applied in sensing devices.

Keywords: tungsten oxide, pH sensor, neural recordings, impedance

1. Introduction

Over the past two decades, the revolution in materials science has driven great advances in all areas of science and engineering. Nanoscience and nanotechnology are leading this revolution

fueled by the industrial progress, the scientific ability to fabricate, model, and manipulate objects with a small number of atoms, and the continuous discovery of new phenomena at the nanoscale [1, 2]. Nanomaterials present unique properties, which are not found in the respective bulk materials [3]. Surface and quantum effects arise in nanostructures due to the large surface-to-volume ratio and to the dimensions that are comparable to the electron wavelength, respectively [4, 5].

In the metal oxides field, the discovery of superconductivity [6] and large magnetoresistance [7] has raised researchers' attention, especially to those with transition metals. Moreover, in traditional electronics, oxides are widely used as semiconductors, dielectrics, and conductive electrodes [8]. In the last years, nanostructured metal oxides for sensing applications have achieved significant advances, mainly due to their better thermal and environmental stability compared with organic materials. These devices, based on nanomaterials, can operate with low power consumption and can be easily integrated with nanoelectronics. Furthermore, the construction of sensors in "low-cost" substrates, such as plastic, paper, or textile, is also in demand for application in portable consumer devices [9–13]. Electrodeposition, in this case, is of great interest due to its flexibility to control the structure and morphology of the oxide electrodes combined with the reduced cost [14, 15].

2. Electrodeposition

The term electrodeposition is often used unclearly, referring either to electroplating or to electrophoretic deposition (EPD) [16]. The electroplating process is based on a solution of ionic species, usually in water, while EPD occurs in a suspension of particles. In electroplating, there is a charge transfer during the deposition to produce the metal or oxide layer in the electrode, while in EPD the deposition occurs without any reaction involved (Fig. 1). In fact, the principal driving force for EPD is the charge and the electrophoretic mobility of the particles in the solvent under the influence of an applied electric field, with the drawback that the solvent should be organic in order to avoid water electrolysis [16, 17].

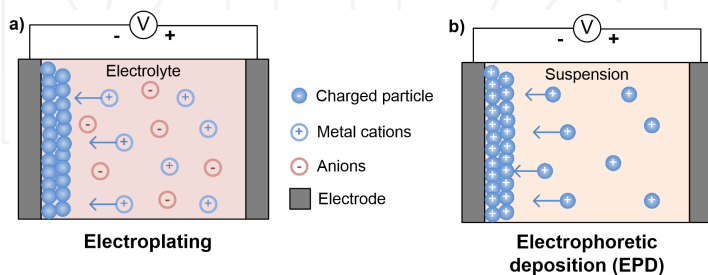


Figure 1. Schematic representation of the two types of cathodic electrodeposition processes: (a) electroplating and (b) electrophoretic deposition (EPD).

Another variation of an electrochemical deposition is the electroless (autocatalytic) deposition in which a reducing agent, dissolved in the electrolyte, is the electron source for the redox reaction, and no external power supply is needed [18]. Nevertheless, the electroless deposition will not be discussed in this chapter.

The first reports on the electrodeposition technique date back to the 19th century; however, the understanding of the process and the electrochemistry involved was only developed in the 20th century and it is believed that further research is still needed to optimize the process [16].

In electroplating, the relation between the current and the overpotential of electrodeposition is given by the Tafel equation (Equation 1), which describes the exponential dependence between the two parameters. Worth mentioning is that with the increase of the overpotential, the ionic current that the electrolyte can supply is limited either by material transport or electrical conductivity [15, 19].

$$i = -FkC \exp\left(\frac{\alpha F \eta}{RT}\right) \quad (1)$$

where i is the current, F is Faraday's constant, k a constant, C the concentration of metal ions in solution (which can be initially dissolved in the electrolyte or originated from the dissolution of the metallic anode), α the coefficient of symmetry (~ 0.5), η the overpotential, R the ideal gas constant, and T the absolute temperature (K).

The first attempt to correlate the amount of particles deposited by EPD with the different parameters influencing electrophoresis was first described by Hamaker for electrophoretic cells with a planar geometry. Over the years, Hamaker's law has been adapted and more recently Equation 2 was derived, relating the weight (W) of the charged particles deposited per unit area of electrode in the initial period, with different parameters, and disregarding the charge of the free ions [20].

$$W = \frac{2}{3} C \varepsilon_0 \varepsilon_r \xi \mu^{-1} E L^{-1} t \quad (2)$$

Here, C is the concentration of the particles; ε_0 and ε_r the permittivity of vacuum and solvent, respectively; ξ the zeta potential of the particles; μ the viscosity of the solvent; E the applied potential; L the distance between the electrodes; and t the deposition time. Equation 2 demonstrates that the deposition weight of the charged particles under ideal EPD depends on all the previous parameters. However, if the solvent, the particles, and the apparatus for EPD are not changed, the weight of the deposited particles (W) is a function of C , E , and t . Therefore, the mass of the deposited particles, namely the thickness of the films, can be easily controlled by the concentration of the suspension, applied potential, and deposition time [17].

Electrodeposition of conventional metals for coatings has a very long history, with more than 200 years for some metals and alloys. Today, electrodeposition is much more than just a

technique for coatings fabrication. In addition to applications such as decorative, wear, and corrosion-protective coatings, electrodeposition is also used for the manufacture of molds, functional coatings for magnetic and electronic applications, and microelectromechanical system components production [5]. In the future, though traditional applications will continue, new ones will rapidly develop, especially in the fields of nanoelectronics, biotechnology, and energy engineering. The electrodeposition of non-metallic materials will become more important and the combination of electrodeposition with other processes will lead to nanostructured materials with new and improved properties [21, 22]. Electrodeposition is extremely versatile and different applications will keep being explored [23].

3. Metal oxide electrodeposition

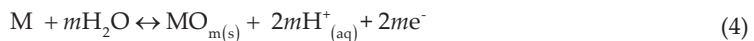
Metal oxides are an important class of materials, which benefit from the large electronegativity of oxygen to induce strong bonding with nearby atoms [22]. At the same time, when compared with bulk materials, nanostructured metal oxides benefit from the spatial confinement, the large fraction of surface atoms, high surface energy, strong surface adsorption, and increased surface-to-volume ratio that greatly improves the performance of these materials [24].

The deposition of nanostructured metal oxides has been already reported by both physical and chemical methods [8, 5]. The advantages of electrodeposition include its speed, low cost, high purity, industrial applicability, use of different types of substrates, and production of films with different morphologies and compositions, as multilayers and alloys [21, 22].

In the electroplating of metal oxides, the reaction involved is usually defined by two consecutive steps (Equation 3). First, the hydroxide will precipitate in the surface of the electrode due to the reaction of the metal ion (M^{n+}) in an alkaline solution, and secondly, the oxide is formed through a condensation/dehydration process. This last step can occur either during electrodeposition or by a subsequent annealing procedure [15].



Another alternative is the formation of metal oxides by anodic oxidation [15]. In this case, the source of the metal ions is the metallic anode and the metal oxide film will be deposited on top of the metal electrode. The general equation can be described as (Equation 4):



In EPD, the metal oxide nanoparticles are generally synthesized by different solution based techniques, e.g., sol-gel, precipitation, and hydrothermal synthesis, prior to deposition. The main challenge of this technique is the preparation of a stable dispersion that originates a film with good properties, uniformity, and appropriate thickness. The use of dispersants, binders,

or other additives that influences the agglomeration and charge of the particles contributes to the tuning of the properties of the deposited film and need to be considered in defining the EPD parameters [16].

3.1. Applications

Nowadays, electrodeposited nanostructured metal oxides are generally used for different applications in laboratories and industry [15]. The latest published reports on the field, listed in Table 1, evidenciate the diversity of areas where these materials can be applied, as presented below.

The deposition of metal/metal oxide nanoparticles composites allowed advances on the protective coatings field. Sajjadnejad et al. [25] improved the corrosion resistance of zinc by co-depositing TiO_2 nanoparticles, while Zeng et al. [26] incorporated CeO_2 nanoparticles to improve the corrosion behavior of nickel coatings. Charlot et al. [27] opened the discussion of the kinetics and mechanism of the anodic EPD of SiO_2 nanoparticles to improve the control of the thickness and properties of these coatings (Fig. 2).

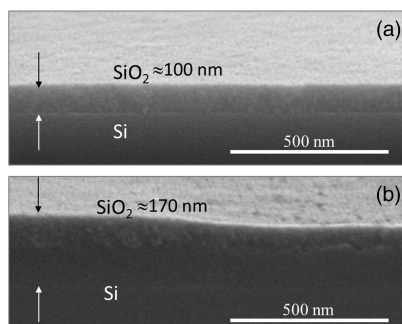


Figure 2. Scanning electron microscopy (SEM) images of a film cross-section obtained from a suspension with a mass fraction of 3% of nanoparticles under an applied electric field of (a) 6 V cm^{-1} and of (b) 60 V cm^{-1} . Reprinted from [27], with permission from Elsevier.

Metal oxide nanostructures are already known to show good catalytic properties. Tu et al. [28] produced Cu_2O -Cu nanoparticles in carbon paper via electroplating. This procedure is an easy, one-step technique that can be an attractive candidate as a visible-light-driven photocatalyst. At the same time, Yoon et al. [29] studied the influence of 2D and 3D structures on electrodeposited Cu_2O films by controlling the electrolyte pH and by using polystyrene (PS) beads as template, respectively. This techniques allowed the production of electrodes with increased surface area. Battaglia et al. [30] also improved the catalytic performance of different Ni electrodes by electrodepositing IrO_2 nanostructures through different electrochemical methods. The composites obtained by galvanostatic deposition of the oxide catalyst presented the best activity for water splitting applications.

Solid oxide fuel cells (SOFC) have shown to be a good alternative for electric power generation systems. SOFC show high energy conversion efficiency, clean power generation, reliability, modularity, fuel adaptability, noise-free, excellent long-term stability, and versatility for direct conversion of chemical energy to electrical energy. In this field, Das and Basu [31] applied the EPD technique to deposit yttria-stabilized zirconia (YSZ) nanoparticles on a NiO-YSZ substrate, which after sintering was suitable for application in SOFC (Fig. 3).

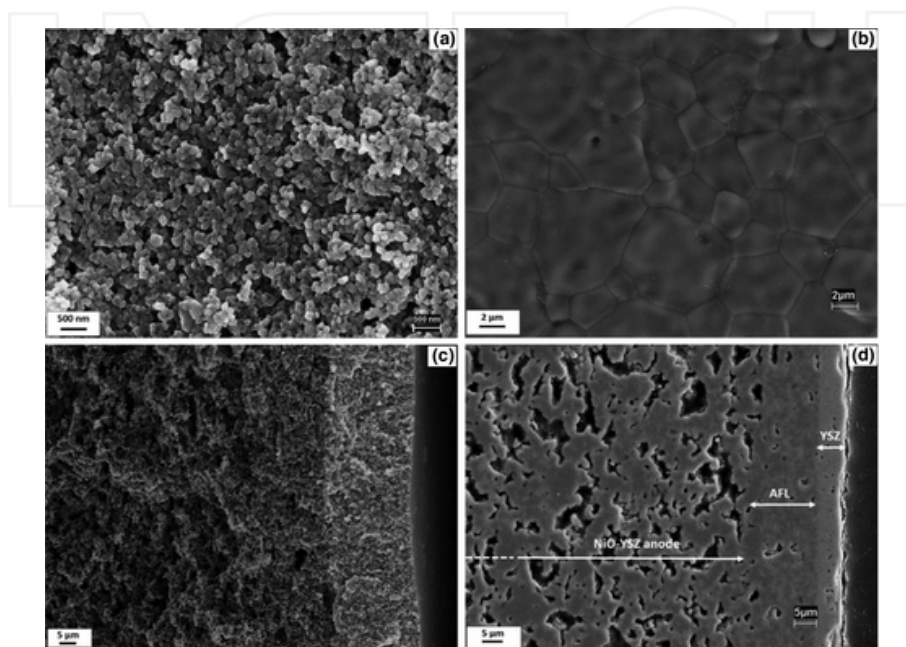


Figure 3. Field-emission SEM images of top view and cross-section of yttria-stabilized zirconia (YSZ) electrophoretic deposition coating (a) (c) as-deposited and (b) (d) sintered at 1400 °C for 6 h, directly deposited onto the conducting polymers such as polypyrrole-coated NiO-YSZ substrate at a constant applied voltage of 15 V. Reprinted from [31], with permission from John Wiley and Sons.

EPD was also the technique used to deposit TiO_2 nanoparticles for dye-sensitized solar cells (DSSC) [32] and Li-ion micro-batteries applications [33]. For DSSC, the thickness of the TiO_2 films was controlled by changing the deposition time and the I_2 dosage that electrically charge the nanoparticles, while for batteries, the EPD was performed with different TiO_2 structures and different 3D aluminum collectors configurations (Fig. 4). The effect of the substrate was also tested in the EPD of ZnO nanoparticles for conductive fabrics applications [34]. Liu et al. [35] studied the EDP of metal oxides using celestine blue as charging and dispersing agent. The nanostructured MnO_2 films were applied for energy storage in electrochemical supercapacitors with high capacitance and excellent capacitance retention at high charge-discharge rates.

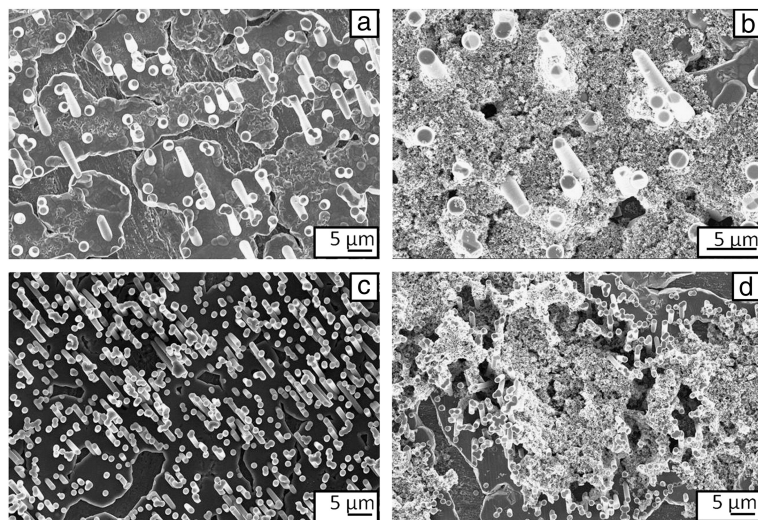


Figure 4. SEM micrographs of aluminium rods obtained by pulsed galvanostatic deposition using a PC membrane with (a) 2 μm -pore size and (c) 1 μm -pore size with electrophoretically deposited P25 particles attached to the respective 3D Al substrates (b) and (d). Reprinted from [33], with the permission from Elsevier.

The use of metal oxides offers functionalities that vary from electrically conducting to insulating and from highly catalytic to inert, which are useful for sensing applications. Different types of metal oxide sensors have been investigated for several decades, and it has been proved that the reduction of crystallite size provided a significant increase in the sensing performances. Even if less established, these type of sensors are very promising and new developments are being accomplished every day [36].

Recently, Cu_2O nanostructures were electroplated to produce a facile and economic photo-electrochemical sensor [37], while Ir_2O_3 was deposited in stretchable and multiplexed pH sensors [38]. This sensor combined electrochemical, microfabrication, and printing techniques and was successfully applied in beating explanted cardiac tissue, with accurate spatiotemporal monitoring of changes in pH (Fig. 5).

Monitoring analgesic drugs with the use of biosensors allows a rapid, reliable, and sensitive method without the requirement of a sample pre-treatment. For that, alloys deposition allows the combination of different materials properties without compromising thickness or surface area available. The biosensors developed by Narang et al. [39] were produced by EPD of an Fe_2O_3 magnetic nanoparticle coated with ZrO suspension containing chitosan, prior to enzyme (horseradish peroxidase) immobilization. Also the combination of Fe_2O_3 with carbon nanotubes and chitosan was earlier used by Batra et al. [40] to immobilize hemoglobin and were applied as an amperometric biosensor.

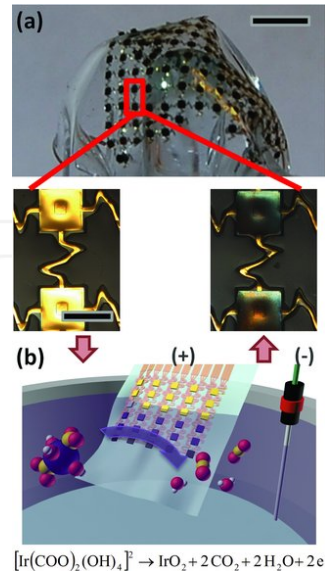


Figure 5. (a) Picture of the produced pH sensors with the magnified images of the gold electrodes before (lower left) and after (lower right) IrO_x electroplating. The scale bars correspond to 5 and 0.5 mm for the upper and the lower images, respectively. (b) Schematic illustration of the chemical reactions during IrO_x electroplating. Reprinted from [38], with permission from John Wiley and Sons.

Application	Nanomaterials/Composites	References
Corrosion and wear resistive coatings	Zn-TiO ₂ , Ni-CeO ₂ , SiO ₂	[25][26][27]
Photocatalyst	Cu ₂ O-Cu, Cu ₂ O	[28][29]
Water splitting	Ni-IrO ₂	[30]
Solid oxide fuel cell	Y ₂ O ₃ -ZrO ₂ (YSZ)	[31]
Dye-sensitized solar cell	TiO ₂	[32]
Li-ion micro-battery	TiO ₂	[33]
Conductive fabric	ZnO	[34]
Supercapacitor	MnO ₂	[35]
Photoelectrochemical sensor	Cu ₂ O	[37]
pH sensor	Ir ₂ O ₃	[38]
Biosensor	ZrO@Fe ₃ O ₄ , cMWCNT-Fe ₃ O ₄	[39][40]

Table 1. List of the latest published research on electrodeposited metal oxide nanostructures/nanomaterials.

4. Nanostructured WO₃

Tungsten oxide (WO₃) is a well-studied semiconductor used for several applications such as chromogenic material, sensor, and catalyst [41]. The major advantages is its low cost and availability, improved stability, reversible change of conductivity and optical properties, high sensitivity, selectivity, and biocompatibility [42].

Transition-metal oxides, especially those with d⁰ and d¹⁰ electronic configurations, as WO₃, TiO₂, or ZnO show interesting properties and stability that are important for sensing applications [43]. The energy band gap of WO₃ corresponds to the difference between the energy levels of the valence band formed by the filled O 2p orbitals and the conduction band formed by empty W 5d orbitals, ranging from 2.6 to 3.25 eV [44]. In nanostructured WO₃, the bandgap generally increases with the reduction of the grain size, which is attributed to the quantum confinement effect [45]. Tungsten oxide is also well known for its properties in a non-stoichiometric form, since its lattice can support a significant concentration of oxygen vacancies [44].

4.1. WO₃ electrodeposition

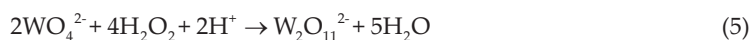
Many liquid and vapor phase synthesis methods have been used to synthesize WO₃ [45]. Nevertheless, for the electrodeposition of nanostructured WO₃ films, more than one method can be adopted: electroplating from a precursor solution [46, 47], anodic oxidation from a metal layer [47–49], and electrodeposition from a WO₃ nanoparticles dispersion [50, 51]. A list of the latest reports is presented in Table 2.

WO ₃ Precursor	Nanostructured film	Application	References
Na ₂ WO ₄	WO ₃	--	[59]
Na ₂ WO ₄	Pt-WO ₃	Proton exchange membrane fuel cell	[53]
Na ₂ WO ₄	TiO ₂ -WO ₃	Photocatalyst	[60]
Na ₂ WO ₄	TiO ₂ -WO ₃	Water splitting	[61]
H ₂ WO ₄	WO ₃ /PANI	Supercapacitor	[52]
PTA	WO ₃	Electrochromic film	[46]
PTA	WO ₃ /PEDOT	Electrochromic film	[62]
W	WO ₃	--	[57]
W	WO ₃ /PANI	Electrocatalyst	[48]
W	TiO ₂ -WO ₃	Photoelectrocatalyst	[63]
W	WO ₃	Photoelectrocatalyst	[64]
W	WO ₃	Photocatalyst	[65]
W	NH ₄ -doped WO ₃	Water splitting	[49]
W	TiO ₂ -WO ₃	Water splitting	[66]

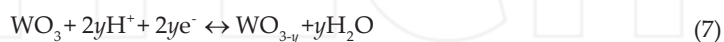
WO ₃ Precursor	Nanostructured film	Application	References
W	WO ₃ – TiO ₂	Electrochromic film	[54]
W	WO ₃	H ₂ sensor	[67]
WO ₃ NPs	WO ₃ /henna	Dye sensitized solar cells	[68]
WO ₃ NPs	WO ₃	Water splitting	[51]
WO ₃ NWs	WO ₃	Electrochromic film	[58]
WO ₃ NPs	WO ₃	Electrochromic film	[50]
WO ₃ NPs	WO ₃	pH sensor	[12]
WO ₃ NPs	WO ₃	Neural electrodes	[69]

Table 2. Resume of the latest published research on electrodeposited nanostructured WO₃ with the respective precursors and final applications.

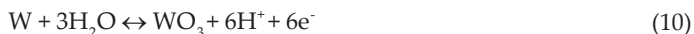
Cathodic electroplating is usually based on the local increase of the pH near the electrode surface due to the reduction of O₂ or H₂O, which induces precipitation of metal ions present in the solution as metal oxide or hydroxide. For the deposition of WO₃, the reactions involved in the formation of the oxide are usually based on the formation of the peroxytungstate (W₂O₁₁²⁻) intermediate from a tungstate salt (or from the reaction of metallic tungsten with hydrogen peroxide), as described in Equations 5 and 6 [46, 52].



Depending on the electrochemical potential and solution pH, the WO₃ phase may also be involved in other reactions, as the formation of sub-stoichiometric oxide and tungsten bronze (Equations 7 and 8) or even re-dissolution of the oxide phase (Equation 9). The reduced phases formed by these reactions have higher conductivity and hydrophilicity than WO₃ and should be considered during characterization of the deposited films [53].



For the anodic oxidation procedure, the general equation can be expressed as Equation 10 [54] and the full mechanism is explained by the occurrence of different reactions simultaneously, as the synthesis of surface oxide films (e.g., W₂O₅, WO₂) and tungstate ions (WO₄²⁻) [55, 56]. This oxidation is usually followed by the slow dissolution of the oxide phase, as in Equations 11 or 12 depending on the solution pH [55, 57].



In the case of the deposition from WO₃ nanoparticles dispersions (EPD), the mechanism is not yet fully understood. The majority of the authors agree that the deposition occurs through an electrophoretic mechanism driven by the surface charge of the particles [51, 58], but in fact, the potential (or current) applied during deposition can also promote tungsten reduction from W⁶⁺ to W⁵⁺ that is counterbalanced by the cation intercalation into the oxide structure, as described in Equation 8, thus forming tungsten bronze (H_xWO₃) [50]. In the work of Liu et al. [50], XRD and optical characterization showed that HWO₃ was obtained as the main phase of the deposited films, which supports the hypothesis of the mechanism via electrochemical reduction. Furthermore, since the reduced WO₃ is significantly more conductive than the oxidized form, it allowed continuous film growth. In the future, further analysis of the deposited films should be conducted to confirm the electrochemical deposition mechanism.

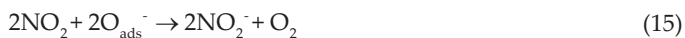
4.2. WO₃ sensing applications

4.2.1. Gas sensors

Precise and affordable monitoring of chemical gases is a critical issue for human health, industrial processes, and environmental protection. For that, nanostructured WO₃ has been intensively studied due to its excellent sensing capabilities and reproducibility. These characteristics are mainly ascribed to the increased surface area and complete depletion of carriers within the nanostructure when exposed to the target gas [45]. The gas sensing mechanism is described by the increase or decrease of the conductance of the oxide layer when exposed to reducing (H₂, H₂S, CO) or oxidizing (NO₂, O₃, CO₂) gases, respectively.



In Equation 13, H₂ adsorbs and reacts with O⁻ formed on the surface of the electrode, increasing the surface conductance and releasing the captured electrons [67].



When NO_2 is targeted on the WO_3 surface, it not only reacts with the electrons from the conduction band (Equation 14) but also with the chemisorbed oxygen (Equation 15), thus promoting a depletion on the surface of the electrode and, consequently, the increase on resistance [70, 71].

An example of a hydrogen gas sensor was built by Yang et al. [67] through anodic oxidation of a tungsten layer previously deposited by radio frequency magnetron sputtering on a sapphire substrate (Fig. 6). The nanoporous WO_3 film sensor, after annealing at 600°C , exhibited good sensitivity to H_2 gas in air.

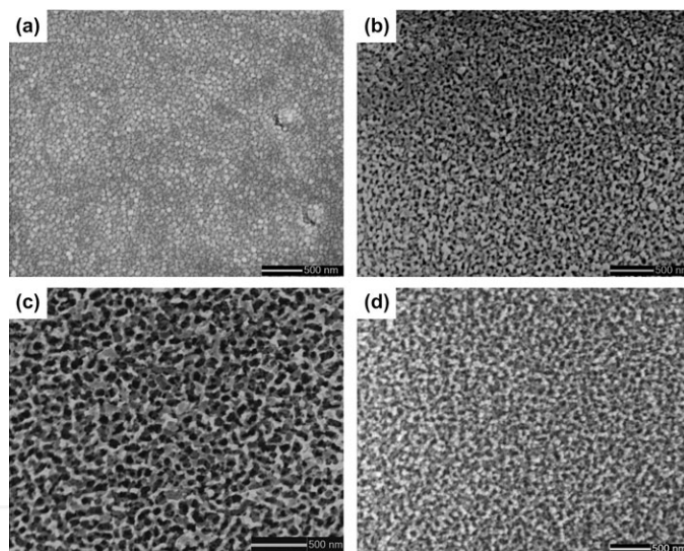


Figure 6. SEM images of tungsten oxide films with different anodic oxidation voltages: (a) 20 V, (b) 30 V, (c) 50 V, and (d) 60 V operating at an electrode distance of 2 cm for 60 min. Reprint from [67], with permission from Cambridge University Press.

4.2.2. Biosensors

The application of WO_3 to other sensing platforms, as in biosensors, is mainly due to the electrical and optical properties mentioned above [72]. In fact, it was already demonstrated that nanoparticles of metal oxides applied to suitable electrode surfaces allow protein immobilization and biocatalytic processes to be driven electrochemically [73]. However, to the best of the authors' knowledge, only Feng et al. [74] employed electrodeposited nanostructured

WO₃ films to enhance the hemoglobin protein loadings, accelerate interfacial electron transfer, and improve thermal stability of the adsorbed protein. The influence of the electrodeposition time to the response time and peak current of the electrode is demonstrated in Fig. 7.

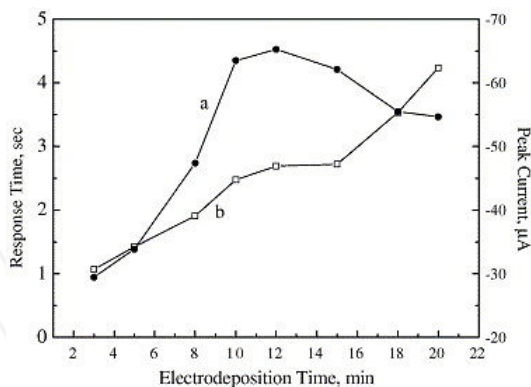


Figure 7. Influence of electrodeposition time on (a) peak current of the cyclic voltammograms in phosphate buffer solution (PBS, pH 6.0) at 100 mV s⁻¹ and (b) typical steady-state response time of Hb/meso-WO₃/graphite electrodes. Reprinted from [74], with permission from Elsevier.

4.2.3. pH sensors

The pH value can be used as an indicator for disease diagnostics, medical treatment optimization, and monitoring of biochemical and biological processes [75]. Nevertheless, the integration of pH sensing systems into the next generation of wearable devices requires a different architecture than currently used in typical glass-type electrodes and a minimal electrode size [76]. In addition, technological and industrial efforts are under way to incorporate different sensors into our daily life by assembling these sensors on common substrates such as plastic, textile, and paper [9]. In the work reported earlier [12], flexible pH sensors were based on electrodeposited WO₃ sensing layer in a gold/polyimide substrate (Fig. 8). The pH sensing mechanism for this material, even if not fully understood, is believed to be dependent of the redox reaction involving the production of the tungsten bronze with a higher conductivity than the tungsten oxide (Equation 8).

4.2.4. Neural electrodes

Microtechnology allowed the arrangement of multiple microelectrodes on the same substrate over small distances (Fig. 9a). Nevertheless, in order to provide sufficient recording sensitivity to small electrodes for measuring neuron electrical activity, they are often coated with different nanostructured or conducting materials to increase the effective surface area and electrochemical interface capacitance [77–79]. The interest in utilizing transition metal oxide films is due to its pseudocapacitive character related to chemisorption processes and redox reactions that take place at the surface [80]. Since nanostructured WO₃ has already proved to enhance capacitive performances due to its large surface area and low charge transport resistance [52],

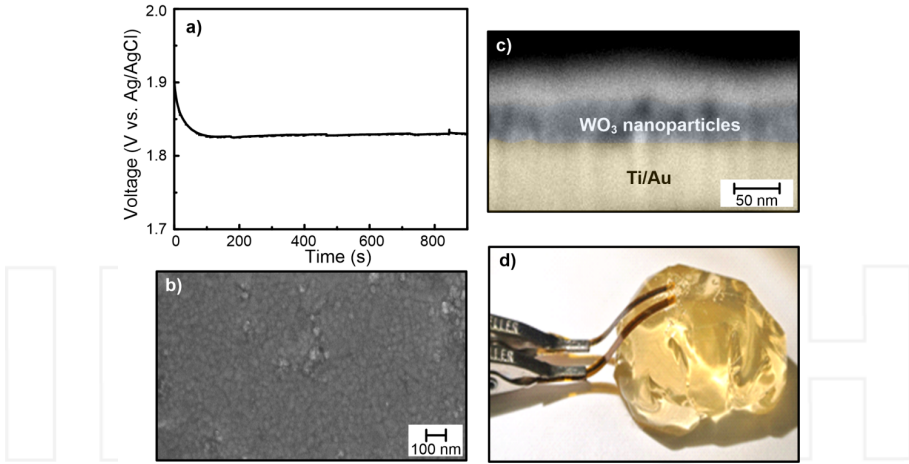


Figure 8. (a) Voltage response during electrodeposition at 20 μA ; (b) topographic and (c) cross-section SEM images of the WO_3 electrodeposited layer; and (d) photograph of the prototype WO_3 sensor using a flexible Ag/AgCl reference electrode in a non-planar surface made of gelatin-based electrolyte [12].

it was used for neural recordings applications [69]. The optimization of the electrodeposition parameters led to a slight increase on the charge storage capacity ($\sim 10\%$) and a decrease of the impedance values, of approximately 40% (Fig. 9b and 9c).

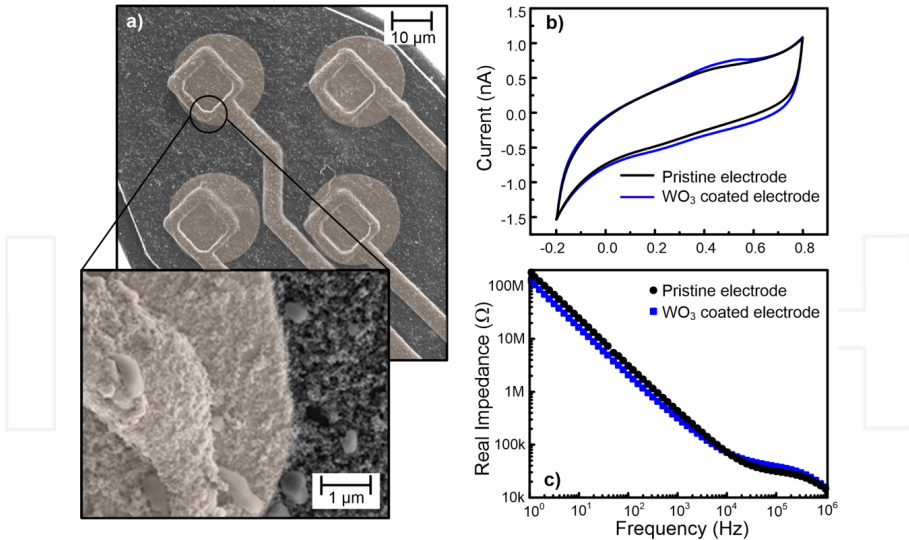


Figure 9. (a) SEM images of the Neuronexus electrode and a detail of the iridium electrode (lighter area) coated with WO_3 nanoparticles, electrodeposited at 30 nA for 15 s; (b) cyclic voltammetry and (c) electrochemical impedance characterizations of the pristine (black) and coated electrodes (blue) [69].

These preliminary results show the versatility of electrodeposition in different materials and configurations as well as in different sensing mechanisms.

5. Conclusions

Tungsten oxide (WO₃) is one of the most studied metal oxide and the sensing performance of this material is of great interest due to the capability of reversible change of both its optical and electrical properties. The evolution in the fields of nanoscience and nanotechnology allowed these materials to replace many organic and metallic materials in a huge range of applications besides creating new areas of development. The increased surface area and the quantum confinement effects in size ranges below 100 nm make nanostructured WO₃ a good platform for gas and pH sensors, along with neural electrodes and biosensors.

In the last decade, the use of electrodeposition for nanostructured metal oxide films has been growing due to the versatility of this method in different applications and materials. Just in the last year, applications varied from catalysts and sensors to capacitors. The use of different types of templates and the deposition of composites will contribute to the continued development of this technique.

Acknowledgements

The authors would like to thank the Portuguese Science Foundation (FCT-MEC) through project EXCL/CTM-NAN/0201/2012, Strategic Project UID/CTM/500025/2013, and doctoral grants SFRH/BD/73810/2010 given to L. Santos and SFRH/BD/76004/2011 given to J. Neto. The authors would like to thank Dr. Adam Kampff from Champalimaud Center of Unknown for the knowledge transfer related with neural electrodes.

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